

**Combined Gas Law
CER**

Name _____
Hour _____

A sample of Carbon dioxide gas occupies 638 mL at 0.893 atm and 12 °C. What will the pressure be at a volume of 881 mL and a temperature of 18 °C?

C = What had a bigger impact on pressure: temperature or volume?

E = Based on your claim, show how the arrows work in relation to your answer from the problem.

R = Discuss what temperature and volume actually measure and associate your claim with the correct gas law.

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